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tr I Name Date: Hour:
Introduction Questions
Draw the Lewis

structures for the
following molecules.

Part c is done for you.

a) H_2O (oxygen is
central) b) SO_2 c) N_2 It
t-t (a [Jrt @ $N=N$ @ @ f\

@ 2 In question 1

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above you should see
three different kinds
of bonds of bonds as
best you can. a) single
bond: one es.iv (ee-)
shave a b) *-l; double
bond:

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to help you complete

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to write down the
definitions and take
down the notes. Circle
and write down any
questions you have so
you do ...

Unit 7: Lewis Structures Chemquest

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Name: _____ Date:

_____ Hour: _____

Information: Structure
of the Atom Note the

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following symbols:

(they are not to scale)

= proton (positive
charge) = electron

(negative charge) =
neutron (no charge)

The following three
diagrams are hydrogen
atoms: 1 1H 2 1H 3 1H

The following three
diagrams are carbon
atoms: Notice the type
of notation used for
atoms: 12 13 14 6C 6 C
6C (6 ...

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prevalent, but answering providers for compact companies are rapidly growing in popularity.

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included in the Lewis Structure. $4 + 3(6) + 2 = 24$ (carbon has four and each oxygen has six; add two for the -2 charge) $5 + 3(1) = 8$ (nitrogen has five;

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each hydrogen has one) Step #2: Draw the “skeleton structure” with the central atoms and the other atoms, each connected with a single bond. Step #3: Add six more electron dots to each atom except the

ChemQuest 23 : Drawing Covalent Compounds

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questions.

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Pathways (first half
videos on ionic
compounds) Video
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- Unit 4 Bonding Basics
& Unit 5 Ionic Bonds
(all videos in each unit)
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**Atomic Structure -
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Draw Lewis structures for covalent compounds. The following procedure can be used to construct Lewis electron structures for more complex molecules and ions:
How-to: Constructing Lewis electron

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structures. 1.

Determine the total number of valence electrons in the molecule or ion. Add together the valence electrons from each atom.

10.5: Writing Lewis Structures for Covalent Compounds ...

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Seichi Morimura "

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